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a curved arrow (\curvearrowright) to

show the movement

of an electron pair. A

sigma(σ) bond is

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formed between atoms by the overlap of two atomic orbitals along the line that connects the atoms. Carbon uses sp^3 -hybridized orbitals to form four such bonds. These bonds are directed from the carbon nucleus toward the corners of a tetrahedron.

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idea almost always

builds on another that

has gone before. It is

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that you keep up with,

or better yet, be a little

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should try to stay one

day ahead of your

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structures Draw

molecule, show all

valence electrons,

add electrons to give

non-hydrogen atoms

full octets Formal

charge = (Valence

electrons) -

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(nonbonding
electrons) - ($\frac{1}{2}$
bonding electrons)

Nonpolar when
electrons are shared
evenly More eneg in
top right (More eneg,
more attraction
towards other atoms)

Molecular dipole -
Sum of individual
dipoles 1.3 Bond
length decreases left
to ...

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